

*Introduction To Chemical Engineering Thermodynamics 7th
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Introduction To Chemical Engineering Thermodynamics

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A few important comments about this version of the international edition. We use this text book in my introductory chemical engineering thermodynamics class and there are other students in my class who also bought the international edition and the chapters are incomplete, there are no English units, and the printing quality is very poor (sometimes with whole sections of tables missing!).

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►example: chemical equilibrium says diamonds should be graphite (or CO₂ if in air)! Sometimes, mixing & reaction are sufficiently fast that equilibrium assumptions are “good enough.”! • if we assume that we can use thermodynamics to obtain the state of the system, we avoid a lot of complexity.! ►do it if it is “good enough!”!

Chemical Engineering Thermodynamics - Sutherland

INTRODUCTION TO CHEMICAL ENGINEERING THERMODYNAMICS J.M. Smith H.C. Van Ness M.M. Abbott LIST OF CORRECTIONS TO FIRST PRINTING OF 6TH EDITION • Page 17, Pbs. 1.7 and 1.8: The atmospheric-pressure values should be 101.78 kPa and 29.86(in Hg), respectively.

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